

Copernicium

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Copernicium is a chemical element with symbol **Cn** and atomic number 112. It is an extremely radioactive synthetic element that can only be created in a laboratory. The most stable known isotope, copernicium-285, has a half-life of approximately 29 seconds. Copernicium was first created in 1996 by the GSI Helmholtz Centre for Heavy Ion Research near Darmstadt, Germany. It is named after the astronomer Nicolaus Copernicus.

In the periodic table of the elements, it is a d-block transactinide element. During reactions with gold, it has been shown^[9] to be an extremely volatile metal and a group 12 element, so much so that it is probably a gas at standard temperature and pressure. Copernicium is calculated to have several properties that differ between it and its lighter homologues, zinc, cadmium and mercury; due to relativistic effects, it may even give up its 6d electrons instead of its 7s ones. Copernicium has also been calculated to possibly show the oxidation state +4, while mercury shows it in only one compound of disputed existence and zinc and cadmium do not show it at all, although more recent calculations cast doubt on this possibility. It has also been predicted to be more difficult to oxidize copernicium from its neutral state than the other group 12 elements.

Isotopes

Copernicium has no stable or naturally occurring isotopes. Several radioactive isotopes have been synthesized in the laboratory, either by fusing two atoms or by observing the decay of heavier elements. Six different isotopes have been reported with atomic masses from 281 to 285, and 277. Most of these decay predominantly through alpha decay, but some undergo spontaneous fission.^[28]

The isotope copernicium-283 was instrumental in the confirmation of the discoveries of the elements flerovium and livermorium.^[30]

Half-lives

Copernicium, ¹¹²Cn

General properties	
Name, symbol	copernicium, Cn
Copernicium in the periodic table	
Atomic number (<i>Z</i>)	112
Group, block	group 12, d-block
Period	period 7
Element category	<div>  transition metal, alternatively considered a post-transition metal</div>
Standard atomic weight (<i>A</i>_r)	[285]
Electron configuration	<div>[Rn] 5f¹⁴ 6d¹⁰ 7s² (<i>predicted</i>)^[1]</div>
<div>per shell</div>	<div>2, 8, 18, 32, 32, 18, 2 (<i>predicted</i>)</div>
Physical properties	
Phase	gas ^[2] (<i>predicted</i>)
Boiling point	<div>357⁺¹¹²_{−108} K (84⁺¹¹²_{−108} °C, 183⁺²⁰²_{−194} °F)^[3]</div>
Density near r.t.	<div>23.7 g/cm³ (<i>predicted</i>)^[1]</div>
Atomic properties	
Oxidation states	<div>2, (1), 0 (parenthesized)</div>

All copernicium isotopes are extremely unstable and radioactive; in general, heavier isotopes are more stable than the lighter. The most stable isotope, ²⁸⁵Cn, has a half-life of 29 seconds. Other isotopes have half-lives shorter than 0.1 seconds. ²⁸¹Cn and ²⁸⁴Cn both have half-lives of 97 ms, and the other two isotopes have half-lives slightly under one millisecond.^[28] It is predicted that the heavy isotopes ²⁹¹Cn and ²⁹³Cn may have half-lives of around 1200 years, and may have been produced in the r-process and be detectable in cosmic rays, though they would be about 10^{−12} times as abundant as lead.^[31]

The lightest isotopes were synthesized by direct fusion between two lighter nuclei and as decay products (except for ²⁷⁷Cn, which is not known to be a decay product), while the heavier isotopes are only known to be produced by decay of heavier nuclei. The heaviest isotope produced by direct fusion is ²⁸³Cn; the two heavier isotopes, ²⁸⁴Cn and ²⁸⁵Cn have only been observed as decay products of elements with larger atomic numbers.^[28] In 1999, American scientists at the University of California, Berkeley, announced that they had succeeded in synthesizing three atoms of ²⁹³Og.^[32] These parent nuclei were reported to have successively emitted three alpha particles to form copernicium-281 nuclei, which were claimed to have undergone an alpha decay, emitting an alpha particle with decay energy of 10.68 MeV and half-life 0.90 ms, but their claim was retracted in 2001.^[33] The isotope, however, was produced in 2010 by the same team. The new data contradicted the previous (fabricated)^[34] data.^[35]

Predicted properties

Chemical

Copernicium is the last member of the 6d series and is the heaviest group 12 element in the periodic table, below zinc, cadmium and mercury. It is predicted to differ significantly from the lighter group 12 elements. The valence s-subshells of the group 12 elements and period 7 elements are expected to be relativistically contracted most strongly at copernicium. This and the closed-shell configuration of copernicium result in it probably being a very noble metal. Its metallic bonds should also be very weak,

oxidation states are predictions)^{[1][4][5]}

1st: 1154.9 kJ/mol

2nd: 2170.0 kJ/mol

3rd: 3164.7 kJ/mol

(more) *(all estimated)*^[1]

empirical: 147 pm


(predicted)^{[1][5]}

122 pm *(predicted)*^[6]

Miscellanea

hexagonal close-packed (hcp)

(predicted)^[7]



54084-26-3

History

after Nicolaus Copernicus

Gesellschaft für Schwerionenforschung (1996)

Most stable isotopes of copernicium

iso	NA	half-life	DM	DE (MeV)	DP
²⁸⁵ Cn	syn	29 s	α	9.15, 9.03?	²⁸¹ Ds
²⁸³ Cn	syn	4 s ^[8]	90% α	9.53, 9.32, 8.94	²⁷⁹ Ds
			10% SF		

possibly making it extremely volatile, like the noble gases, and potentially making it gaseous at room temperature.^{[1][36]} However, it should be able to form metal-metal bonds with copper, palladium, platinum, silver, and gold; these bonds are predicted to be only about 15–20 kJ/mol weaker than the analogous bonds with mercury.^[1]

Once copernicium is ionized, its chemistry may present several differences from those of zinc, cadmium, and mercury. Due to the stabilization of 7s electronic orbitals and destabilization of 6d ones caused by relativistic effects, Cn^{2+} is likely to have a $[\text{Rn}]5f^{14}6d^87s^2$ electronic configuration, using the 6d orbitals before the 7s one, unlike its homologues. The fact that the 6d electrons participate more readily in chemical bonding means that once copernicium is ionized, it may behave more like a transition metal than its lighter homologues, especially in the possible +4 oxidation state. In aqueous solutions, copernicium may form the +2 and perhaps +4 oxidation states.^[1] The diatomic ion Hg_2^{2+} , featuring mercury in the +1 oxidation state, is well-known, but the Cn_2^{2+} ion is predicted to be unstable or even non-existent.^[1] Copernicium(II) fluoride, CnF_2 , should be more unstable than the analogous mercury compound, mercury(II) fluoride (HgF_2), and may even decompose spontaneously into its constituent elements. In polar solvents, copernicium is predicted to preferentially form the CnF_5^- and CnF_3^- anions rather than the analogous neutral fluorides (CnF_4 and CnF_2 , respectively), although the analogous bromide or iodide ions may be more stable towards hydrolysis in aqueous solution. The anions CnCl_4^{2-} and CnBr_4^{2-} should also be able to exist in aqueous solution.^[1]

Physical and atomic

Copernicium should be a very heavy metal with a density of around 23.7 g/cm³ in the solid state; in comparison, the most dense known element that has had its density measured, osmium, has a density of only 22.61 g/cm³. This results from copernicium's high atomic weight, the lanthanide and actinide contractions, and relativistic effects, although production of enough copernicium to measure this quantity would be impractical, and the sample would quickly decay.^[1] However, some calculations predict copernicium to be a gas at room temperature, the first gaseous metal in the periodic table^{[1][36]} (the second being flerovium), due to the closed-shell electron configurations of copernicium and flerovium.^[37] The atomic radius of copernicium is expected to be around 147 pm. Due to the relativistic stabilization of the 7s orbital and destabilization of the 6d orbital, the Cn^+ and Cn^{2+} ions are predicted to give up 6d electrons instead of 7s electrons, which is the opposite of the behavior of its lighter homologues.^[1]

In addition to the relativistic contraction and binding of the 7s subshell, the $6d_{5/2}$ orbital is expected to be destabilized due to spin-orbit coupling, making it behave similarly to the 7s orbital in terms of size, shape, and energy. Hence copernicium may not be a noble transition metal, but rather a semiconductor^[3] with a band gap of around 0.2 eV.^[7] Copernicium is expected to

crystallize in the hexagonal close-packed crystal structure, with lattice parameters $a = 332$ pm and $c = 540$ pm. The c/a ratio of 1.63 is the ideal value, establishing a kinship between solid copernicium and the solid noble gases, though its cohesive energy (enthalpy of crystallization) should be on the order of that of mercury rather than be near the lower value of the noble gases.^[7]

Experimental atomic gas phase chemistry

Interest in copernicium's chemistry was sparked by predictions that it would have the largest relativistic effects in the whole of period 7 and group 12.^[1] Copernicium has the ground state electron configuration $[Rn]5f^{14}6d^{10}7s^2$ and thus should belong to group 12 of the periodic table, according to the Aufbau principle. As such, it should behave as the heavier homologue of mercury and form strong binary compounds with noble metals like gold. Experiments probing the reactivity of copernicium have focused on the adsorption of atoms of element 112 onto a gold surface held at varying temperatures, in order to calculate an adsorption enthalpy. Owing to relativistic stabilization of the 7s electrons, copernicium shows radon-like properties. Experiments were performed with the simultaneous formation of mercury and radon radioisotopes, allowing a comparison of adsorption characteristics.^[38]

The first experiments were conducted using the $^{238}\text{U}(^{48}\text{Ca},3n)^{283}\text{Cn}$ reaction. Detection was by spontaneous fission of the claimed parent isotope with half-life of 5 minutes. Analysis of the data indicated that copernicium was more volatile than mercury and had noble gas properties. However, the confusion regarding the synthesis of copernicium-283 has cast some doubt on these experimental results. Given this uncertainty, between April–May 2006 at the JINR, a FLNR–PSI team conducted experiments probing the synthesis of this isotope as a daughter in the nuclear reaction $^{242}\text{Pu}(^{48}\text{Ca},3n)^{287}\text{Fl}$. In this experiment, two atoms of copernicium-283 were unambiguously identified and the adsorption properties indicated that copernicium is a more volatile homologue of mercury, due to formation of a weak metal-metal bond with gold, placing it firmly in group 12.^[38]

In April 2007, this experiment was repeated and a further three atoms of copernicium-283 were positively identified. The adsorption property was confirmed and indicated that copernicium has adsorption properties completely in agreement with being the heaviest member of group 12.^[38] These experiments also allowed the first experimental estimation of copernicium's boiling point: 84_{-108}^{+112} °C, so that it may be a gas at standard conditions.^[3]

Because the lighter group 12 elements often occur as chalcogenide ores, experiments were conducted in 2015 to attempt to deposit copernicium atoms on a selenium surface to form copernicium selenide, CnSe . Reaction of copernicium atoms with trigonal selenium to form a selenide were observed, with $\Delta H_{\text{ads}}^{\text{Cn}}(\text{t-Se}) > 48$ kJ/mol, with the kinetic hindrance towards selenide formation being lower for copernicium than for mercury. This was unexpected as the stability of the group 12 selenides tends to decrease down the group from ZnSe to HgSe , while it increases down the group for the group 14 selenides from GeSe to PbSe .^[39]

Source

- Wikipedia: Copernicium (<https://en.wikipedia.org/wiki/Copernicium>)